

**CHM201 General Chemistry and Laboratory I**  
**Laboratory 7 – Acetic Acid Concentration in Vinegar**  
*CHM201-03/T1/T2*

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Section 03

Lab Group B

You must complete **two** successful titrations of each type (KHP and vinegar). Record the data from **all** titrations and mark the two successful titrations. **Remember to estimate volume to the second decimal place,  $\pm 0.05\text{mL}$ .**

**WARNING: You are NOT recording the volume of NaOH solution in the buret! You are recording the reading from the buret. If your initial volume is recorded as  $\sim 50\text{ mL}$ , YOU DID IT WRONG! If you are unsure about how to read the buret, consult me before you attempt to do it.**

**Results (Sodium Hydroxide Standardization) – You only need TWO good titrations!**

KHP Titrations	1	2	3	4	5
Mass of KHP (4 decimal places)	1.0040g	1.0570g			
Initial Volume $V_i$ NaOH (2 decimal places)	0.01	0.03			
Final Volume $V_f$ NaOH (2 decimal places)	<del>15.01</del> 20.98	21.81			
Successful (Yes or No)	Yes	Yes			

**Results (Acetic Acid Content of Vinegar) – You only need TWO good titrations!**

Vinegar Titrations	1	2	3	4	5
Volume of Vinegar (2 decimal places)	10.0	10.0			
Initial Volume $V_i$ NaOH (2 decimal places)	0.02	0.01			
Final Volume $V_f$ NaOH (2 decimal places)	35.89	35.45			
Successful (Yes or No)	Yes	Yes			